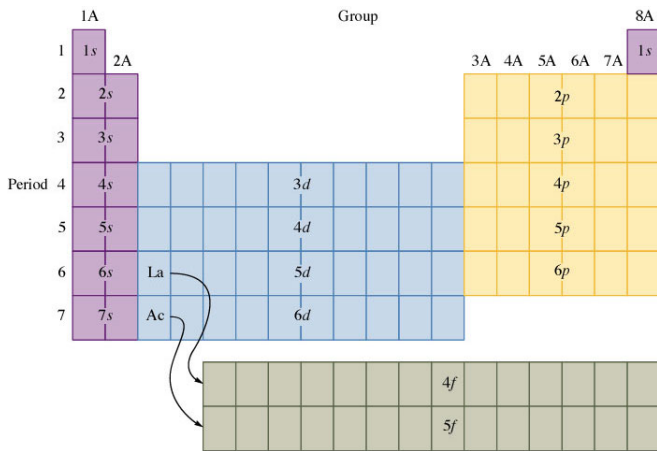
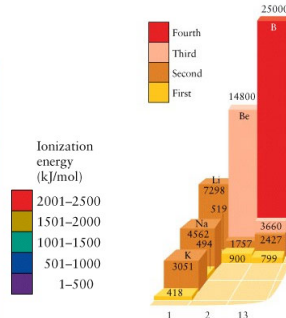
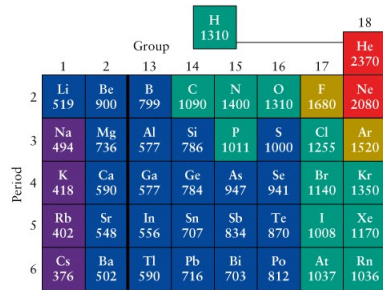
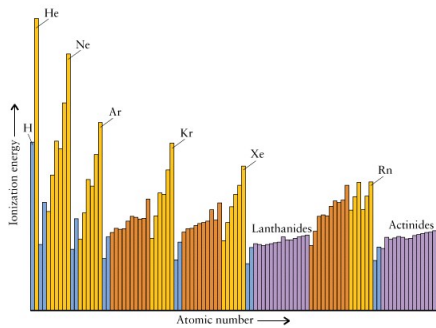


Periodic Trends



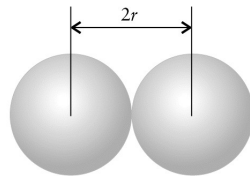
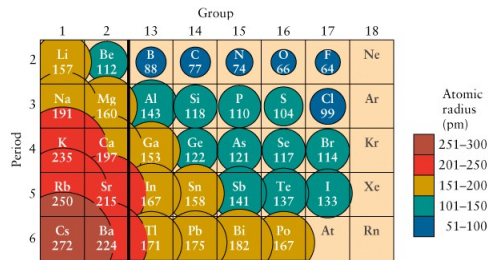
Ionization Energy – The energy change associated with the removal of one electron from a neutral atom <<<+1 Ion is formed>>>

(First) Ionization Energy – How hard is it to pull the concertgoer with the worst seat away from the concert?



Jones, L. and Atkins, P., Chemistry: Molecules, Matter and Change, 4th ed., W. H. Freeman: New York (2000).

Atomic Radius – What is the occupied radius of the concert venue? [Superstar bands have better audio and video stuff]



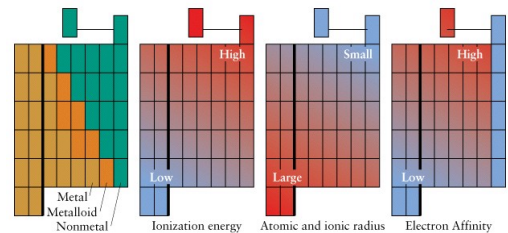
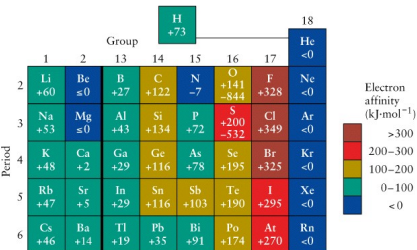
17 Atomic radius

Jones, L. and Atkins, P., Chemistry: Molecules, Matter and Change, 4th ed., W. H. Freeman: New York (2000).

Electron Affinity – The energy change associated with the addition of an extra electron to a neutral atom <<<-1 Ion is formed>>>

Electron Affinity – How desirable is the next available seat?

Summary:



Jones, L. and Atkins, P., Chemistry: Molecules, Matter and Change, 4th ed., W. H. Freeman: New York (2000).

Electronegativity – A measure of the relative ability of an atom to attract and retain electrons. Sort of like the popularity of a movie star in that the more popular they are, the more photographers (electrons) tend to follow them around. Electronegativity takes into account both the ability of an atom to attract new electrons and their ability to hold on to the electrons they have. It is an inherently approximate measure, but the difference in the electronegativity between two atoms serves as a pretty good predictor of how ionic a bond between them will be, and which of them will have the lion's share of the shared electrons.